

Homework

Sodium chloride (NaCl)

- When dissolved in water, it only dissociates into sodium and chloride ions.
- These ions do not react with water, so there is no hydrolysis.

Ammonium bicarbonate (NH_4HCO_3)

- When dissolved, it first dissociates into ammonium ions and bicarbonate ions.
- Then, both ions can undergo hydrolysis:
 - Ammonium ion makes the solution slightly acidic.
 - Bicarbonate ion can act both as a weak acid and a weak base.
- The solution is usually slightly acidic.

Summary:

- Sodium chloride → dissociation only.
- Ammonium bicarbonate → dissociation + hydrolysis.