

Questions ?

Unit 1 | Lesson 4

Answered

Aqueous Solutions and their Properties

Interactive test



First Multiple Choice Questions

Questions signed by * are answered in detail.

- Which of the following expresses water in all water bodies on the Earth's surface?
 - It exists in a pure form.
 - It contains dissolved or suspended substances.
 - It has the same temperature.
 - It has the same degree of salinity.
- Which of the following is from the colligative properties of aqueous solutions?
 - Elevation of vapor pressure.
 - Depression in boiling point.
 - Elevation of freezing point.
 - Osmotic pressure.
- The colligative properties of aqueous solutions depend on
 - nature of solvent
 - nature of solute
 - volume of solution
 - number of solute molecules
- What happens to the vapor pressure of an aqueous solution when increasing the concentration of the solute?
 - It decreases.
 - It increases.
 - It isn't affected.
 - It vanishes.
- The liquid starts to boil when its vapor pressure is
 - less than the pressure exerted on it
 - greater than the pressure exerted on it
 - equal to the pressure exerted on it
 - double the pressure exerted on it
- What happens to the vapor pressure of the liquid when its temperature increases in a closed system?
 - It increases.
 - It decreases.
 - It isn't affected.
 - It vanishes.
- * At the standard atmospheric pressure, if the boiling point of a pure solvent = $X^{\circ}\text{C}$, the boiling point of the solution that produced from dissolving one of its salts in it is
 - equal to $X^{\circ}\text{C}$
 - greater than $X^{\circ}\text{C}$
 - less than $X^{\circ}\text{C}$
 - equal to $\frac{1}{2} X^{\circ}\text{C}$