

November Chemistry Test

Grade: Second Secondary

Question One:

Write what explains the following conclusions from Rutherford's experiment:

1. Most of the atom is empty space and not a solid sphere.

2. There is a part inside the atom that has a large density and occupies a very small space.

3. The charge of the dense part equals the charge of the positively charged particles.

Question Two:

What are the results of the following:

1. The return of the excited electron to its original energy level.

2. The absorption of some different quanta of energy by atoms.

Question Three:

Mention the advantages and disadvantages of Bohr's model.

Question Four:

Calculate the concentration of the resulting solution from mixing:

100 mL of sodium chloride solution of concentration 0.2 M

with 400 mL of sodium chloride solution of concentration 0.1 M.

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